

# *Doon Public School*

A Senior Secondary School  
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**Assignment – 2**

**Subject: chemistry**

**Class: XI**

**Date: 09-08-2018**

## **INSTRUCTIONS:**

- **Do the given assignment in your note book.**

1. Explain why the electron gain enthalpy of fluorine is less negative than that of chlorine.
2. All transition elements are d-block elements, but all d-block elements are not transition elements. explain.
3. Nitrogen has positive electron gain enthalpy whereas oxygen has negative. However, oxygen has lower ionisation enthalpy than nitrogen. explain.
4. Write the drawbacks in Mendeleev's periodic table that led to its modification.
5. How does atomic radius vary in a period and in a group? How do you explain the variation?
6. The first ionization enthalpy of magnesium is higher than that of sodium. On the other hand, the second ionization enthalpy of sodium is very much higher than that of magnesium. Explain.
7. Describe the main characteristic properties of s,p,d,f block elements.